

# Intro to Biochemistry

Principal biological elements:

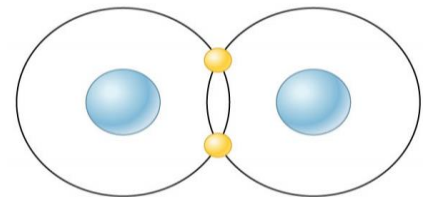
Periodic Table of the Elements

1	IA																O																														
1	H																	2	He																												
3	Li	IIA																		5	B	6	C	7	N	8	O	9	F	10	Ne																
11	Na	12	Mg	IIIB		IVB	VB	VIB	VII	VII		IB		IIB	13	Al	14	Si	15	P	16	S	17	Cl	18	Ar																					
19	K	20	Ca	21	Sc	22	Ti	23	V	24	Cr	25	Mn	26	Fe	27	Co	28	Ni	29	Cu	30	Zn	31	Ga	32	Ge	33	As	34	Se	35	Br	36	Kr												
37	Rb	38	Sr	39	Y	40	Zr	41	Nb	42	Mo	43	Tc	44	Ru	45	Rh	46	Pd	47	Ag	48	Cd	49	In	50	Sn	51	Sb	52	Te	53	I	54	Xe												
55	Cs	56	Ba	57	*La	72	Hf	73	Ta	74	W	75	Re	76	Os	77	Ir	78	Pt	79	Au	80	Hg	81	Tl	82	Pb	83	Bi	84	Po	85	At	86	Rn												
87	Fr	88	Ra	89	+Ac	104	Rf	105	Ha	106	Sg	107	Ns	108	Hs	109	Mt	110	111	112	113																										
				* Lanthanide Series																																											
				58	Ce	59	Pr	60	Nd	61	Pm	62	Sm	63	Eu	64	Gd	65	Tb	66	Dy	67	Ho	68	Er	69	Tm	70	Yb	71	Lu																
				+ Actinide Series																																											
				90	Th	91	Pa	92	U	93	Np	94	Pu	95	Am	96	Cm	97	Bk	98	Cf	99	Es	100	Fm	101	Md	102	No	103	Lr																

Valence electrons:

Reactivity

Covalent bonds



Why are hydrocarbon good candidates for building blocks?